



Sodium Chlorite Neutralization

Introduction

If sodium chlorite is spilled or becomes a waste, it must be disposed of in accordance with local, state, and federal regulations by a NPDES permitted out-fall or in a permitted hazardous waste treatment, storage, and disposal facility.

Due to the reactivity of sodium chlorite, neutralization for disposal purposes should be avoided whenever possible. Where permitted, the preferred method for handling sodium chlorite spills and waste is by dilution, as discussed in the OxyChem Safety Data Sheet (SDS) for sodium chlorite in Section 6, (Accidental Release Measures). Sodium chlorite neutralization procedures must be carried out only by properly trained personnel wearing appropriate protective equipment.

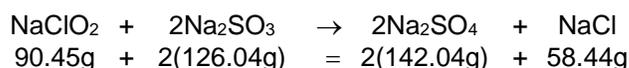
Reaction Considerations

If a specific situation requires sodium chlorite to be neutralized, the chlorite must first be reduced by a reaction with sodium sulfite. The use of sodium sulfite is recommended over other reducing agents such as sodium thiosulfate ($\text{Na}_2\text{S}_2\text{O}_3$), sodium bisulfite (NaHSO_3), and sodium meta-bisulfite ($\text{Na}_2\text{S}_2\text{O}_5$). These other agents may result in acidic solutions and should not be used due to the possibility of generating chlorine dioxide gas.

The reaction of sodium chlorite with sodium sulfite is illustrated below. In this reaction, one mole of sodium chlorite (NaClO_2) reacts with two moles of sodium sulfite (Na_2SO_3) to form two moles of sodium sulfate (Na_2SO_4) and one mole of sodium chloride (NaCl). The equivalent weights of these compounds are listed below the reaction. It is important to note that this reaction

sodium sulfite per pound of active sodium chlorite solution and mix thoroughly. For a diluted (5%) sodium chlorite solution, this is

is exothermic and liberates a significant amount of heat (ΔH).



$$\Delta H = -168 \text{ kcal/mole NaClO}_2$$

For example, when starting with a 5% NaClO_2 solution, the heat generated from this reaction could theoretically raise the temperature of the solution by 81°C (146°F). Adequate dilution, thorough mixing and a slow rate of reaction are important factors in controlling the temperature increase (ΔT).

Procedure

The complete neutralization procedure involves three sequential steps: dilution, chlorite reduction, and alkali neutralization. The dilution step lowers the strength of the sodium chlorite solution to 5% or less; the reduction step reacts the diluted chlorite solution with sodium sulfite to produce a sulfate solution, and the neutralization step reduces the pH of the alkaline sulfate solution from approximately 12 to 4-5.

1. **Dilution:** Due to the potential for heat generation, it is important to dilute the sodium chlorite with water to a solution containing no more than 5% (wt.) sodium chlorite. At 5%, the pH of the solution should be approximately 11.9. When diluting sodium chlorite in a reactor, the reactor should never be filled beyond 75% of its capacity. A reactor that provides good agitation is beneficial.
2. **Chlorite Reduction:** Slowly and carefully, add dry sodium sulfite at a ratio of 2.8 lbs. of dry equivalent to 1.2 lbs. (548 grams) of dry sodium sulfite per gallon of solution (197





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grams of active sodium chlorite). The addition of an excess of sodium sulfite will assure the complete reduction of chlorite and will prevent the generation of chlorine dioxide in the next step. As a precaution, sodium sulfite should be added slowly to allow for adequate heat dissipation.

Thorough mixing of the sodium sulfite is very important to prevent the generation of chlorine dioxide during the addition of hydrochloric acid in step 3.

3. **Alkali Neutralization:** Once the chlorite has been completely reacted, slowly add a 5% hydrochloric acid (HCl) solution to lower the pH to 4-5 and mix thoroughly to complete the neutralization. HCl should be added at a ratio of 1 gallon of 5% HCl per 10 gallons of reduced 5% sodium chlorite from step 2. Monitor and maintain the solution temperature below 100°F. Rapid additions of acid will cause excessive temperature increases. Additions of HCl in excess of the 1/10 ratio will contribute to corrosion problems and safety concerns.

The use of stronger acid concentrations can result in the generation of extreme heat and material splattering and should be avoided.

4. Monitor and maintain the pH of the solution between 4-5. Use acid/base pH test paper with a range of pH 2-10. If too much acid is added, sodium bicarbonate (common baking soda) can be used to raise the pH. Carbon dioxide gas is liberated during the addition of sodium bicarbonate to acid. Appropriate safety precautions may need to be taken if a significant quantity of carbon dioxide gas is generated.
5. **Verify Neutralization is Complete:** Sample the neutralized product in a clear beaker or

equivalent. Maintain a pH of 4-5, monitoring with acid/base pH paper. As a final indicator, test the solution with KI/starch indicator paper. Neutralization is complete when the paper does not turn black, blue, or brown in color.

6. If the paper turns black, blue, or brown, neutralization is not complete and chlorine dioxide could be generated. Add additional sodium sulfite, mix, and repeat steps 4, 5 and 6.
7. Dispose of the neutralized material in accordance with local, state, and federal regulations.

Further Information

More detailed information on sodium chlorite is available upon request through the OxyChem Technical Service Department. Call or write to:

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